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Basic Thermodynamic
Formulas (Exam
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Control Mass (no mass
flow across system

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Thermodynamics

boundaries)

Conservation of mass:

$$\dot{m} = \dot{m}_{in} - \dot{m}_{out}$$

Conservation of energy

(1st Law): $\dot{m} - \dot{m} = \Delta \dot{m} =$

$$\Delta \dot{m} + \Delta \dot{m}_{in} + \Delta \dot{m}_{out} = \dot{m} \Delta \dot{m} +$$

$$\dot{m}^2 - \dot{m}^2 + \dot{m}^2 +$$

$$\dot{m}(\dot{m}^2 - \dot{m}^2)$$

Basic

Thermodynamic Formulas (Exam Equation Sheet)

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Engineering Formula
Sheet - Madison Local
Schools For quasi-static
and reversible
processes, the first law
of thermodynamics is:
$$dU = \delta Q - \delta W$$
$$\{ \displaystyle dU = \delta Q - \delta W \}$$
where δQ is the heat
supplied to the system
and δW is the work
done by the system.

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ME 300:

THERMODYNAMICS II

BASIC EQUATION

SHEET FALL 2020

Governing Principles

Mass Balance: CV ie ie

$\frac{dm}{dt}$ Closed

System: m constant

Energy Balance: $\frac{dE}{dt} = \dot{Q} - \dot{W} + \sum \dot{m}_i h_i - \sum \dot{m}_e h_e + \sum \dot{m}_i gz_i - \sum \dot{m}_e gz_e$

CV ie CV CV i i i e e e

ie dE QW mh gz mh gz

dt

ME 300:

THERMODYNAMICS II

BASIC EQUATION

SHEET FALL 2020 ...

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ME 211 and ME312

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Abata, April 1, 2020

Conservation of mass:
where Boundary work
any system: and flow
work (open system) ,
assuming ideal gas and
since $T=C$ then and For
the polytropic process,
that is : Open system
work: , ,

ME 211 and ME312
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General equation .

Valid at any instance of
timeSteady or not

steady flow. Usually

Simplifies to $Q C.V m$

$ih i W C.V . m eh m 2u$

" $m 1u 1$ steady

state. $m i m e q w h e"$

$i h e h i m 1 m 2$ State

1 State 2 Second law

Non- flow $m s 2" s 1 -Q$

$T S gen s 2" s 1 -q T s$

gen steady transient 0

$m is i" m es e -Q T S$

gen 0 $s i " e -q T s gen$

$m 2s 2 " m 1s 1 m is i "$

$m es e -q T s gen$

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This list gives you some of the most common conversion factors you need in thermodynamics.

Acceleration: $1 \text{ m/s}^2 = 100 \text{ cm/s}^2$. Area: $1 \text{ m}^2 = 10^4 \text{ cm}^2 = 10^6 \text{ mm}^2$. Density: $1 \text{ g/cm}^3 = 1 \text{ kg/L} = 1,000 \text{ kg/m}^3$. Energy, heat, work, internal energy, enthalpy: $1 \text{ kJ} = 1,000 \text{ J}$

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$$= 1,000 \text{ N}\cdot\text{m} = 1 \text{ kPa}\cdot\text{m}$$

$$3. 1 \text{ kJ/kg} = 1,000 \text{ m}^2 / \text{s}^2.$$

**Thermodynamics For
Dummies Cheat
Sheet - dummies**

Thermodynamics ΔT A

$v = A^2 v P =$ rate of

heat transfer $Q =$

thermal energy $A =$

Area of thermal

conductivity $U =$

coefficient of heat

conductivity (U-factor)

$T =$ change in

temperature $R =$

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resistance to heat flow
(R-value) $v =$ velocity
net = net power
radiated = $5.6696 \times 10^{-8} \sigma T^4$
 T_1 = temperature at time 1,
time 2 $v =$ flow velocity

Engineering Formula Sheet

Purdue's School of Mechanical Engineering conducts world-class research in robotics, automotive, manufacturing, rocket and jet propulsion,

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nanotechnology, and
much more. ME 200 -
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**Thermodynamics I -
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$S = k_B (\ln \Omega)$
{\displaystyle S=k_{
{B}} (\ln \Omega)} ,
where k_B is the
Boltzmann constant,
and Ω denotes the

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volume of macrostate
in the phase space or
otherwise called
thermodynamic
probability. $dS = \delta Q$
 T . $\{\displaystyle dS =$
 $\{\frac{\{\delta Q\}}$
 $\{T\}\}$, for reversible
processes only.

Table of thermodynamic equations - Wikipedia

All of the thermodynamics
in one sheet Figure 1:
thermodynamics.

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Figure 2: polytropic process diagrams.

Figure 3: first and second laws diagrams.

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or Pa) Absolute
Temperature, T ($^{\circ}\text{R}$ or
 K) Volume, V (ft^3 or m^3)
 v $v_{\text{m}} = (\text{ft}^3/\text{lbm}$ or
 $\text{m}^3/\text{kg})$ Internal Energy,
 U (Btu or kJ) u $u_{\text{m}} =$
(usually in Btu/lbm or
 kJ/kg) Enthalpy, H (Btu
or kJ) Enthalpy, $h = u$

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+ $Pv = H/m$ (usually in
Btu/lbm or kJ/kg)
Entropy, S (Btu/°R or
kJ/K)

FE Reference

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Engineering**

Thermodynamics key
facts (7/9) • Ideal gas
law • 1. st. form :

$Pv = \frac{RT}{M}$ • P = Pressure, v = Volume,

P = Pressure,

v = Volume,

n = number of

molecules, R = gas constant.

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= Boltzmann's constant,

T = temperature [in K]

• 2. nd. form : $U = nRT$

n = number of moles, R =

Revision :

Thermodynamics

For quasi-static processes where changes in kinetic and potential energy are

not important. $dU = Q - pdV$

$du = q - pdv$

$dH = Q + Vdp$ or $dh = q + vdp$

+ vdp

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**COMPENDIUM OF
EQUATIONS Unified
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Thermodynamics is filled with equations and formulas. Here's a list of the most important ones you need to do the calculations necessary for solving thermodynamics problems. Combustion equations: Air-fuel ratio: Hydrocarbon fuel

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combustion reaction:

Compressibility

calculations:

Compressibility factor

Z: $Pv = ZRT$ Reduced

temperature: Reduced

pressure: Pseudo-

reduced specific

volume ...

Important Thermodynamic Equations and Formulas - dummies

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sounds intimidating
and it can

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THERMODYNAMICS In order to accurately and precisely discuss various aspects of thermodynamics, it is essential to have a well-defined vernacular. As such, a list of some foundational concepts and their definitions are shown

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