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405 1. Calculate the ratio of effusion rates for nitrogen (N_2) and neon (Ne). $\text{Rate } N_2 / \text{Rate } Ne = 0.849$

$$\frac{\text{Rate } N_2}{\text{Rate } Ne} = \sqrt{\frac{\text{molar mass } Ne}{\text{molar mass } N_2}} = \frac{20.18}{28.02} = 0.849$$

2. Calculate the ratio of diffusion rates for carbon monoxide and carbon dioxide. $\text{Rate } CO / \text{Rate } CO_2 = 1.25$

$$\frac{\text{Rate } CO}{\text{Rate } CO_2} = \sqrt{\frac{\text{molar mass } CO_2}{\text{molar mass } CO}} = 1.25$$

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4.2 For the reaction, 2A

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the rate = $k[A][B]^2$

with $k = 2.0 \times 10^{-6}$

$\text{mol}^{-2} \text{L}^2 \text{s}^{-1}$. Calculate

the initial rate of the

reaction when $[A] =$

0.1 mol L^{-1} , $[B] = 0.2$

mol L^{-1} . Calculate the

rate of reaction after

$[A]$ is reduced to 0.06

mol L^{-1} . Sol. Initial rate

of reaction = $k[A][B]^2$

$= (2.0 \times 10^{-6} \text{ mol}^{-2}$

$\text{s}^{-1}) (0.1 \text{ mol L}^{-1}) (0.2$

$\text{mol L}^{-1})^2 = 8 \times$

$10^{-9} \text{ mol L}^{-1} \text{ s}^{-1}$.

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CH9 CH10 CH11 CH12

CH13 CH14 CH15 CH16

CH17 CH18 CH19 CH20

CH21 CH22 CH23 CH24

CH25 CH26 CH27

Problem: 1P 2P 3P 4P

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12P 13P 14P 15P 16P

17P 18P 19P 20P 21P

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22P 23P 24P 25P 26P

27P 28P 29P 30P 31P

32P 33P 34P 35P 36P

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Chapter 12. 26 terms.

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terms. Agriculture

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CH4 CH5 CH6 CH7 CH8
CH9 CH10 CH11 CH12
CH13 CH14 CH15 CH16
CH17 CH18 CH19 CH20
CH21 CH22 CH23 CH24
CH25 CH26 CH27 CH28
CH29 CH30 CH31

Problem: 1P 2P 3P 4P
5P 6P 7P 8P 9P 10P 11P
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211 CHAPTER 11
SOLUTIONS MANUAL

Section 11.1

Assessment page 372

5. Compare the mass of the reactants and the mass of the products in a chemical reaction, and explain how these masses are related. The coefficients in the balanced equation indicate the molar relationship between

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