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15: Chemical Equilibrium. We introduced

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the concept of equilibrium in Chapter 11, where you learned that a liquid and a vapor are in equilibrium when the number of molecules evaporating from the surface of the liquid per unit time is the same as the number of molecules condensing from the vapor phase. Vapor pressure is an example of a physical equilibrium because only the physical

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form of the substance changes.

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3/16/11 15-3 As the reaction proceeds, the product concentration increases and the reverse reaction begins to take place. $\text{FeSCN}^{2+}(\text{aq}) \rightleftharpoons \text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq})$ Rate (reverse) = $k_{\text{reverse}} [\text{FeSCN}^{2+}]$ As the reaction continues, the rate of the forward reaction decreases (because $[\text{Fe}^{3+}]$ and

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of this section, you will be able to:

Describe the nature of equilibrium systems; Explain the dynamic nature of a chemical equilibrium;

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